

Solubility of Ionic Compounds in Water			
Negative Ions (anions)	+	Positive Ions (cations)	form Compounds which are:
All		Group 1 ions (Li^+ , Na^+ , K^+ , Rb^+ , etc.)	Soluble
All		Ammonium ion, NH_4^+	Soluble
Nitrate, NO_3^-		All	Soluble
Acetate, $\text{C}_2\text{H}_3\text{O}_2^-$		All	Soluble
Chloride, Cl^- Bromide, Br^- Iodide, I^-		Ag^+ , Pb^{2+} , Hg_2^{2+} , Cu^+ All Others	Not Soluble Soluble
Sulfate, SO_4^{2-}		Ca^{2+} , Ba^{2+} , Sr^{2+} , Pb^{2+} All Others	Not Soluble Soluble
Sulfide, S^{2-}		Group 1 ions, NH_4^+ , Group 2 ions (Be^{2+} , Mg^{2+} , etc.) All Others	Soluble Soluble Insoluble
Hydroxide, OH^-		Group 1 ions, NH_4^+ , Ba^{2+} , Sr^{2+} All others	Soluble Not Soluble
Phosphate, PO_4^{3-} Carbonate, CO_3^{2-} Sulfite, SO_3^{2-}		Group 1 ions, NH_4^+ All others	Soluble Not Soluble
<u>Note:</u> Soluble means more than 0.1 mole will dissolve per liter			